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hscn knowing the number of moles of  $\text{Fe}^{3+}$  and hscn present in the equilibrium mixture and the volume of the mixture we can easily find the concentrations of those two species mol  $\text{Fe}^{3+}$   $1.70 \times 10^{-5}$  mol  $\text{Fe}^{3+}$   $8.50 \times 10^{-4}$  M

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