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Rachel Sandford

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solvent water n_a 54 g 18 g mol 1 3 mol total number of
moles $n_a + n_b$ 0 5 3 3 5 mol mole fraction of solute ethyl
alcohol $x_b = \frac{n_b}{n_a + n_b} = \frac{0.5}{0.5 + 3} = 0.1429$

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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