

# El Cautivo De Jorge Luis Borges

E Durkheim

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methanol  $CH_3OH$  and water in a solution prepared by dissolving 45 g of alcohol in 40 g of  $H_2O$  molar mass of  $H_2O$  is 18gm mole and molar mass of  $CH_3OH$  is 32gm mole solution moles of  $CH_3OH$   $\frac{45}{32} = 1.4$  mole moles of  $H_2O$   $\frac{40}{18} = 2.2$  moles therefore according to the equation

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 02 11 11 g of urea nh<sub>2</sub>conh<sub>2</sub> was  
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 the molarity and molality of the solution  
 given n 14 h 1 c 12 o 16 given mass of  
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 water 100 g 0 1 kg to find molarity of  
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 c 6 h 14 d 0 660 g cm 3 when it is  
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 cm<sup>3</sup> in a 35 0 55 0 ml ratio assume that  
 the volumes are additive

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 no units on it and is written as a  
 decimal value do not change it to  
 percent note of caution you could see  
 the term mole percent it is simply the  
 mole fraction multiplied by 100 for  
 example in the problem just below the  
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 its mole percent would be 28 85  
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