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period there are two types of nuclear
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63 2 15 0 30 0 65 30 9 29 2 2 12 6 15
32 30

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137 55 cs 137 56 ba x because the
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of the equation x must have a

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aq h 2 g li s n 2 g li 3 n s ca oh 2 hbr
cabr 2 h 2 o c 4 h 10 o 2 co 2 h 2 o nh 3
cuo cu n 2 h 2 o q10 5 5 balance the

following equations
 $\text{Fe} + \text{Cl}_2 \rightarrow \text{FeCl}_3$
 $4\text{H}_2\text{O} + 2\text{CO} \rightarrow 2\text{H}_2\text{O} + \text{NaOH} + 3\text{AsO}_3\text{H}$

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web identify the missing particle in the following nuclear reaction
 $1\text{ }^0_0\text{n} + 235\text{ }^{92}_{92}\text{U} \rightarrow 137\text{ }^{52}_{42}\text{Te} + 97\text{ }^{38}_{38}\text{Sr} + 40\text{ }^{97}_{40}\text{Zr}$

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 bismuth 212 decays into polonium 212 beryllium 8 and a positron are produced by the decay of an unstable nucleus neptunium 239 forms from the reaction of uranium 238 with a neutron and then spontaneously converts into plutonium 239 strontium

90 decays into

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web answer 4 correct in the reactants and products must be the same part b what is the value of in the following nuclear reaction
 $6\text{ }^5_9\text{I} \rightarrow \text{ }^6_9\text{I} + \text{ }^0_0\text{n}$
 express your answer as an integer hint 1 how to approach the problem use rule 2 the total of the subscripts atomic numbers mathematically this can be written as answer

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