

Of Stars And Tides

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 alcohol $x_b = \frac{n_b}{n_a + n_b} = \frac{0.5}{0.5 + 3.5} = 0.125$

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 by kilograms solute $0.449 \text{ mol kg}^{-1} \times 2.92 \text{ kg} = 1.31108 \text{ mol}$ of
 kbr 2 moles times molar mass equals grams $1.31108 \text{ mol} \times 119.0023 \text{ g mol}^{-1} = 156 \text{ g}$ kbr

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 molarity of hexane C_6H_{14} $\text{d} = 0.660 \text{ g cm}^{-3}$ when it is mixed
 with heptane C_7H_{16} $\text{d} = 0.684 \text{ g cm}^{-3}$ in a 35.0 : 55.0 ml ratio
 assume that the volumes are additive

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web notice that the mole fraction has no units on it and is
 written as a decimal value do not change it to percent note of
 caution you could see the term mole percent it is simply the
 mole fraction multiplied by 100 for example in the problem
 just below the mole fraction of cinnamic acid is 0.2885 its
 mole percent would be 28.85

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