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 states and write the oxidation half
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 and for each half reaction balance the
 atom undergoing redox changes if
 necessary add the number of electrons
 that correspond to the change in
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 oxidation is loss of e reduction is gain
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 al b ba2 c br 2 d ca e ga3 f h 2 g h

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 oxidation numbers of atoms change if
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 number increases ie it becomes more
 ve or less ve if an atom is reduced its
 oxidation number decreases ie it
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 oxidation state of each element in the
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 ch 3 oh 2 identify the oxidation state of
 nitrogen in the following a no 3 b no 2 c
 no 2

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 reaction is one in which one or more
 atoms change oxidation numbers
 oxidation occurs when an atom s
 oxidation state becomes more positive
 indicating that electrons have been lost
 reduction occurs when an atom s
 oxidation state becomes less positive
 indicating that electrons have been
 gained

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 commonly known as redox reactions
 are reactions that involve the transfer

of electrons from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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reactions for each redox reaction the first one is done for you 23 $\text{Fe}^{2+} \rightarrow \text{Co}^{2+}$ $\text{Fe} \rightarrow \text{Fe}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} \rightarrow \text{Ni}^{3+}$ 3 $\text{Ag} \rightarrow \text{Ag}^{+} + \text{e}^-$ oxidation reduction 25 $\text{Cu}^{2+} \rightarrow \text{Pb}^{2+}$ $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^-$ oxidation reduction 26 $\text{O}_2 \rightarrow \text{O}^{2-}$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{sr} + \text{o}_2 \rightarrow 2\text{sro}$ sr^0 to sr^{2+} oxidized reducing agent o^0 to o^{2-} reduced ox ag $2\text{li} + \text{s} \rightarrow \text{li}_2\text{s}$ li^0 to li^{+} oxidized red ag s^0 to s^{2-} reduced ox ag $3\text{cs} + \text{br}_2$

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