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## R Sandford

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web 1 consider the reaction  $\text{i}_2\text{O}_5 + 5 \text{CO} \rightarrow 5 \text{CO}_2 + \text{I}_2$  a 80.0 grams of iodine pentoxide ( $\text{i}_2\text{O}_5$ ) reacts with 28.0 grams of carbon monoxide ( $\text{CO}$ ) determine the mass of iodine ( $\text{I}_2$ ) which could be produced b if in the above situation only 0.160

moles of iodine  $I_2$  was produced i what mass of iodine was produced

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web limiting reactant theoretical and percentage yields key 4  $K_2O$  s 2 h 2 o l 4 koh aq 3  $O_2$  g if a reaction vessel contains 0 mol  $K_2O$  and 0 mol h 2 o what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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the balanced equation shown below if the reaction of 90 6 grams of co produces 36 7 grams of  $C_3H_8$  what is the percent yield  $3CO + 7H_2 \rightarrow C_3H_8 + 2O$

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web start with 8 3 mole  $H_2S$  and multiply it with 2 mole  $H_2O$  over 2 mole  $H_2S$  from the equation  $8 3 \text{ mole } H_2S \times 2 \text{ mole } H_2O / 2 \text{ mole } H_2S \times 18g \text{ } H_2O / 1 \text{ mole } H_2O = 149 4g \text{ } H_2O$  is your theoretical yield so percent yield  $137 1g \text{ } H_2O / 149 4g \text{ } H_2O \times 100 = 91 77\%$  problems 6 8 are similar to problem 5

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this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$   
 4  $\text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$   
 160.0 g  $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 / 302.3 \text{ g Al}_2\text{O}_3 = 26.982 \text{ g Al}_2\text{O}_3$   
 4 mol  $\text{Al} \times 1 \text{ mol Al}_2\text{O}_3$  c determine the percent yield  
 $\text{yield} = \frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{26.0}{26.982} \times 100 = 86.01\%$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6  
 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13  
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web the percent yield is calculated as follows  
 $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$   
 based on this definition we would expect a percent yield

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 $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$  calculate the theoretical yield of  $\text{CH}_3\text{OH}$  if 68.5 g of  $\text{CO}$  is reacted with 8.6 g of  $\text{H}_2$   
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 $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4$   
 which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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consider the reaction  $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$  80.0 grams of iodine v oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of carbon monoxide  $\text{CO}$  determine the mass of iodine  $\text{I}_2$  which could be produced 80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  xs 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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