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SJ Ball

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 0.444 mol of CoCl_2 in 0.654 L of solution
 98.0 g of phosphoric acid H_3PO_4 in 1.00 L of solution
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 $141.9579 \text{ g/mol} \times 0.0129 \text{ m} = 2 \text{ molality}$
 $0.825 \text{ g} / 141.9579 \text{ g/mol} = 0.00581158 \text{ mol}$
 $0.00581158 \text{ mol} / 0.4500 \text{ kg} = 0.0129 \text{ m}$
 3 mole fraction $n_2 = \frac{0.825 \text{ g}}{141.9579 \text{ g/mol}} = 0.00581158 \text{ mol}$
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