

Korea National University Of Arts

Rosina Ehmann

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In a global used by displays and the ceaseless chatter of instantaneous connection, the melodic elegance and emotional symphony created by the published term frequently disappear into the background, eclipsed by the persistent sound and

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web teachers in general believe that grammar is central to language learning and students need direct and explicit teaching of grammar rules for accuracy burges ethe

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web teachers beliefs about grammar teaching can determine whether they prioritize explicit instruction focus on communicative approaches or adopt a combination of both these

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web this qualitative study attempts to explore secondary level experienced esl english as a second language teachers beliefs concerning the teaching of grammar employing

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web oxidation reduction redox reactions oxidation $\text{Cu} \rightarrow \text{Cu}^{2+} + 2\text{e}^-$ reduction $\text{MnO}_4^- + 4\text{H}^+ + 3\text{e}^- \rightarrow \text{MnO}_2 + 2\text{H}_2\text{O}$ when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^-$ is

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what is the coefficient in front of OH^- in the balanced equation for the following redox reaction in basic solution

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web balance this redox reaction by the half reaction method by assuming an acidic solution $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$ balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution balancing redox reactions in basic

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow \text{AuCN} + 2\text{e}^-$ $2\text{H}_2\text{O}_2 + 2\text{e}^- \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^-$ convert to basic solution $4\text{CN}^- + 2\text{Au} + 2\text{OH}^- \rightarrow 2\text{AuCN} + 2\text{H}_2\text{O} + 2\text{CN}^-$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 + 2\text{OH}^- \rightarrow 2\text{H}_2\text{O} + 2\text{OH}^-$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web a redox equation can be balanced using the following stepwise procedure
 1 divide the equation into two half reactions
 2 balance each half reaction for mass and charge
 3 equalize the number of electrons transferred in each half reaction
 4

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method
 step 1 split reaction into half reactions reduction and oxidation
 step 2 balance the charge or oxidation number with electrons
 step 3 balance ce o by adding ce h₂o

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web bonus problem here is a balanced net ionic equation
 $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_3^- + 3\text{H}_2\text{O}$
 the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0
 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$
 now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1
 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$
 solution 1 balanced half reactions
 $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
 $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$
 2 equalize the electrons
 $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
 $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$
 multiplied by a factor of 6

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web balance each half reaction in basic solution
 $4\text{S} + 6\text{H}_2\text{O} + 7\text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{NO}_3^- + \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}_2\text{O}_3 + \text{NO}_2 + \text{Mn}_2\text{O}_3$
 balance each redox reaction in acid solution using the half reaction method
 $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{TeO}_4 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}_2\text{O}_3 + 10\text{H}^+$
 $11\text{PbO}_2 + \text{IO}_3^- + 4\text{H}^+ \rightarrow \text{PbO}_2 + 12\text{AsO}_3 + \text{IO}_3^- + 2\text{H}_2\text{AsO}_4$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c 6 h 8 o 6 2h 2no 2 c 6 h 6 o 6 2h 2 o 2no choose 1 answer c 6 h 8 o 6 and no 2 a c 6 h 8 o 6 and no 2 no 2 and c 6 h 8 o 6 b no 2 and c 6 h 8 o 6 h and no 2 c h

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