

# Divisioni Con La Prova Del 9

## SJ Ball

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properties

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to create a visual representation of the

following periodic trends atomic size

ionization energy and electronegativity

atomic radius the size of an atom

measured in either nanometers nm

angstroms Å

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web jan 28 2019 by mapping specific

trends directly onto a blank periodic

table the activity uses the periodic table

as a foundational aspect of the major

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connections between placement on the

periodic table and a specific trend

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web you might need periodic table the

first two ionization energies for

beryllium are shown below be g be g e

be g be 2 g e i 1 900 kJ mol i 2 1757 kJ

mol which of the following identifies

the most probable value for the third

ionization energy for be and provides

the best justification

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introduce students to periodic trends it

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properties for all elements in the

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properties of the elements follow a

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lab periodic law states that when the

elements are arranged in order of

increasing atomic number there is a

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and physical properties elements within

the same group have the most similar

chemical and physical properties

although not identical

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vocabulary atomic radius electron

affinity electron cloud energy level

group ion ionization energy metal

nonmetal nucleus period periodic

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increasing atomic radius carbon

aluminum oxygen potassium oxygen

carbon aluminum potassium 2 rank the

following elements by increasing

electronegativity sulfur oxygen neon

aluminum neon aluminum sulfur

oxygen 3 why does fluorine have a

higher ionization energy than iodine

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key background in this investigation

you will examine several periodic

trends including atomic radius

ionization energy and ionic radius you

will

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as you move from left to right across a

period explain your answer in terms of

the number of protons in the nucleus 2

how does atomic radius change as you

move from top to bottom in a group

explain your answer in terms of the

number of electron shells 3 on the

representation of the periodic table

below



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