

Oh Virgen La Ms Hermosa Letra

S Marginson

Question #22e51 - Socratic Products formed are; $\text{Cu}(\text{OH})_2 + \text{HCl} \rightarrow \text{CuCl}_2 + \text{H}_2\text{O}$ Since CuCl_2 is soluble in water.. But; $\text{Cu}(\text{OH})_2 + \text{HCl} \rightarrow \text{CuCl}_2 + \text{H}_2\text{O}$ " [Not Balanced]" Balancing the equation.. *Question #18488 - Socratic* The degree of dissociation $\alpha=0.0158$ $K_b=2.51 \times 10^{-6}$ color (white) (x)"mol/l") Triethylamine is a weak base and ionises: $\text{CH}_3)_3\text{N} + \text{H}_2\text{O} \rightleftharpoons (\text{CH}_3)_3\text{N}^+(\text{H}) + \text{OH}^-$ For which: $K_b = \frac{[(\text{CH}_3)_3\text{N}^+(\text{H})][\text{OH}^-]}{[(\text{CH}_3)_3\text{N}]}$ Rearranging and taking -ve logs of both sides we get the ... *Question #44ec4 - Socratic* That's a silly way to say carbon dioxide! $\text{CO}_2(\text{g}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{CO}_3(\text{aq})$ May be wrong, but it seems reasonable. The gas would interact with the hydroxide and water molecules to form carbonic acid. There are also concentrations of OH^- and H^+ leftover from this reaction and from the auto-ionization of water, but that's probably going too far into detail. What are the spectator ions in the reaction between sodium Oct 16, 2016 · The nitrate and the sodium ions. $\text{Na}_2\text{CO}_3(\text{aq}) + 2\text{AgNO}_3(\text{aq}) \rightleftharpoons \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NaNO}_3(\text{aq})$ The net ionic equation is: $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightleftharpoons \text{Ag}_2\text{CO}_3(\text{s})$... **What is the product of the following reaction? 1) CH_3OH - Socratic** Dec 23, 2017 · These are ostensibly acid-base reactions.... For ammonium we could write.... $\text{NH}_4^+ + \text{HO}^- \rightleftharpoons \text{NH}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$ For methanol, the acid base reaction would proceed ... **Question #71ce2 - Socratic** Nov 9, 2014 · $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ when the acid was added to the resulting solution. The H^+ and OH^- react in a 1:1 ratio. This tells us that the number of moles of H^+ used will be equal to the number of OH^- moles in solution. Likewise, 2 moles of lithium produces 2 moles of OH^- . This is also a 1:1 ratio. *Could a buffered solution be made by mixing aqueous ... - Socratic* Mar 25, 2017 · Could a buffered solution be made by mixing aqueous solutions of HCl and NaOH? Explain. Why isn't a mixture of a strong acid and its conjugate base considered a buffered solution? **In each of the following pairs, which one has the lower viscosity?** Jul 31, 2015 · $\text{C}_2\text{H}_5\text{Cl}$ has weak dipole-dipole forces from the C-Cl bonds, but $\text{CH}_3\text{-OH}$ has strong hydrogen bonds. Confirmation: The viscosity of $\text{C}_2\text{H}_5\text{Cl}$ is 0.279 mPa·s, while the viscosity of $\text{CH}_3\text{-OH}$ is 0.593 mPa·s. *Question #627ef - Socratic* functional groups like -OH, -CHO, -CO- are present in carbohydrates. commonly two functional are present in a carbohydrate, -OH and -CHO, or -OH and -CO-. Aldose shows different reactions and test like Tollen's silver mirror test, felling's test. different types of reactions are also shown by carbohydrates like Osazone synthesis, length increasing reaction, ... *What is the chemical formula for lithium hydroxide? - Socratic* Apr 15, 2017 · Lithium is a Group 1 metal and commonly

forms a M^+ ion. Hydroxide anion, OH^- , has a unit negative charge. When they make music together, there is thus 1:1 stoichiometry between ions: $Li(s) + H_2O(l) \rightarrow LiOH(aq) + \frac{1}{2} H_2(g) \uparrow$ □ □

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