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present in the given space

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to express relative amounts of

substances in a mixture is with the

mole fraction mole fraction x is the

ratio of moles of one substance in a

mixture to the total number of moles of

all substances for a mixture of two

substances ce a and ce b the mole

fractions of each would be written as

mole fraction percentage by mass

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moles of solute ethyl alcohol n b 23 g

46 g mol 1 0 5 mol number of moles of

solvent water n a 54 g 18 g mol 1 3 mol

total number of moles n a n b 0 5 3 3 5

mol mole fraction of solute ethyl alcohol

x b n b n a n b 0 5 3 5 0 1429

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methanol ch 3 oh and water in a

solution prepared by dissolving 4 5 g of

alcohol in 40 g of h 2 o molar mass of h

2 o is 18gm mole and molar mass of ch

3 oh is 32gm mole solution moles of ch

3 oh 4 5 32 0 14 mole moles of h 2 o 40

18 2 2 moles therefore according to the

equation

13 2 units of concentration chemistry

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web we used mole fractions in gen

chem 1 to describe gas mixtures where

the partial pressure of a gas was its

mole fraction time the total pressure p

a x ap t mole fraction a left frac mole a

moles total right sum n 1 all x n 1 mole

fraction of substance a is often denoted
by x a and

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all the mole fractions is equal to one x a

x b 1 2 the multiplication of the mole

fraction by 100 gives the mole

percentage 3 mole fraction is a unitless

and dimensionless expression

advantages of molar fraction mole

fraction is not dependent on the

temperature

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web nov 6 2023 xa mol a mol a mol b

and xb mol b mol a mol b note in the

given mixture the sum of all the mole

fractions present is equal to one xa xb 1

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water 100 g 0 1 kg to find molarity of solution solution

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web calculate the mass percent mole fraction molality and molarity of hexane C_6H_{14} d 0 660 g cm 3 when it is mixed with heptane C_7H_{12} d 0 684 g

cm³ in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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