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oxide  $\text{I}_2\text{O}_5$  reacts with 28.0 grams of  
carbon monoxide  $\text{CO}$  determine the  
mass of iodine  $\text{I}_2$  which could be  
produced b if in the above situation  
only 0.160 moles of iodine  $\text{I}_2$  was  
produced i what mass of iodine was  
produced

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percentage yields key  $4 \text{KO}_2 + 2 \text{H}_2\text{O} + 4$   
 $\text{KOH} \rightarrow 3 \text{O}_2 + 4 \text{KOH}$  if a reaction vessel  
contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$   
what is the limiting reactant

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quantities of two or more reactants are

given in a stoichiometric problem you  
must identify the click the card to flip  
limiting reagent

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in an amount lower than required by  
the reaction stoichiometry thus limiting  
the amount of product generated  
percent yield measure of the efficiency  
of a reaction expressed as a percentage  
of the theoretical yield

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calculator periodic table given the  
following reaction  $\text{Cu} + 2 \text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{Ag}$

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percentage yield 1 for the balanced  
equation shown below if the reaction of  
90.6 grams of  $\text{CO}$  produces 36.7 grams  
of  $\text{C}_3\text{H}_8$  what is the percent yield  $3\text{CO}$

$7\text{H}_2 + \text{C}_3\text{H}_8 \rightarrow 3\text{H}_2\text{O}$

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web start with 8.3 mole  $\text{H}_2\text{S}$  and  
multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  
 $\text{H}_2\text{S}$  from the equation  $8.3 \text{ mole H}_2\text{S} \times 2$   
 $\text{mole H}_2\text{O} / 2 \text{ mole H}_2\text{S} \times 18 \text{ g H}_2\text{O} / 1 \text{ mole}$   
 $\text{H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$  is your theoretical yield  
so percent yield  $137.1 \text{ g H}_2\text{O} / 149.4 \text{ g H}_2\text{O} \times 100 = 91.77\%$   
problems 6.8  
are similar to problem 5

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and percent yields** - Mar 30 2022

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reagents theoretical actual and percent  
yields example of a limiting reagent  
problem how many grams of  $\text{NH}_3$  can  
be produced theoretically from the  
reaction of 5.0 g of  $\text{N}_2$  what is the  
limiting reagent if 8.52 g are actually  
formed what is the

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web limiting reagents problem one 1  
for the balanced equation  $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO}$   
 $3\text{H}_2\text{O}$  if the reaction of 21.3 grams of  $\text{O}_2$   
produces a 47.3% yield how many grams  
of  $\text{CO}$  would be produced step one

figure out how many moles of the limiting reagent you have o 2 16 2 32 grams mole 21 3 grams 0 66 moles 32 g m

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained calculated mass of the product expected  $\times 100$  in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction al o<sub>2</sub> al<sub>2</sub>O<sub>3</sub> 4 al 3 o<sub>2</sub> 2 al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of al<sub>2</sub>O<sub>3</sub> 160 0 g al  $\times$  1 mol al  $\times$  2 mol al<sub>2</sub>O<sub>3</sub>  $\times$  101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual  $\times$  100 260 0 g  $\times$  100 86 01 theoretical 302 3 g

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86 6 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is

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web the percent yield is calculated as follows text percent yield  $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  percent yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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when the reaction is complete

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sodium phosphate what is the  
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 yield worksheet 1 consider the reaction  
 $\text{i2o5} + 5 \text{co} \rightarrow 5 \text{co2} + \text{i2}$  g a 80 0 grams  
 of iodine v oxide i2o5 reacts with 28 0  
 grams of carbon monoxide co  
 determine the mass of iodine i2 which  
 could be produced 80 g i2o5 1 mol i2o5  
 1 mol i2 xs 1 333 8 g i2o5 1 mol i2o5 28  
 g co 1 mol co

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 reaction is the ratio of the actual yield  
 to the theoretical yield multiplied by  
 100 to give a percentage text percent  
 yield text actual yield g over text  
 theoretical yield g times 100 label 3 7 3

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