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 $2\text{ICl} + \text{H}_2 \rightarrow 2\text{HI} + 2\text{HCl}$
 rate $k[\text{ICl}]^2$
 a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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 b order of the reaction with respect to acetaldehyde is 3/2 and overall order is also 3/2

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3 b \rightarrow 5 c 2 d

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a q h a 2 g the reaction between mg s and hcl a q is
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