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hint s, k_m

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 $2 \text{ICl} + \text{H}_2 \rightarrow 2 \text{HI} + 2 \text{HCl}$
rate $k [\text{ICl}]^2 [\text{H}_2]$
a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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particles must have both the proper amount of energy and
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for the reaction below if substance a is disappearing at a rate
of $1.82 \times 10^{-2} \text{ mol l}^{-1} \text{ s}^{-1}$ at what rate is c appearing
 $3 \text{ a} \rightarrow 5 \text{ c} + 2 \text{ d}$

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law for the overall reaction rate $k_1 \text{ a} + 2 \text{ b} \rightarrow 2 \text{ c}$ consider the
following mechanism
 $\text{a} + 3 \text{ o}_2 \rightarrow \text{fast} + 3 \text{ o}_2$
 $\text{a} + 2 \text{ o}_2 \rightarrow \text{slow}$
a write the overall balanced chemical equation
 $2 \text{ a} + 3 \text{ o}_2 \rightarrow 2 \text{ b}$
identify any intermediates within the mechanism
c what is the order with respect

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d write the rate law for the overall reaction rate $k_1 \text{ a} + 2 \text{ b} \rightarrow 2 \text{ c}$
consider the following mechanism
 $\text{a} + 3 \text{ o}_2 \rightarrow \text{fast} + 3 \text{ o}_2$
 $\text{a} + 2 \text{ o}_2 \rightarrow \text{slow}$
a write the overall balanced chemical equation
 $2 \text{ a} + 3 \text{ o}_2 \rightarrow 2 \text{ b}$
identify any intermediates within the mechanism
c
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represented by the equation above in a kinetics experiment a
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