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 2 balance each half reaction for mass and charge
 3 equalize the number of electrons transferred in each half reaction
 4

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 $3\text{Br}_2 + 6\text{OH}^- \rightarrow 3\text{Br}_2 + 3\text{H}_2\text{O}$
 the above equation takes place in two stages first bromine reacts with oh ions to form bromide ions and bro ions second the bro ions react to form bromide ions and bromate v ions

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 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$
 solution 1 balanced half reactions
 $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
 $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$
 2 equalize the electrons
 $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$
 $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$
 multiplied by a factor of 6

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adding h 2 o h and e in this order balance the atoms in the equation apart from o and h

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