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Yicheng Fang

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 calorimeter th water in it and the hot
 object made up of copper which is put
 in the calorimeter are the same the
 initial temperature of the calorimeter
 and water is 30 c and that of hot object
 is 60 c the specific heat of copper and
 water are 0 09 cal gm c and 1 cal gm c
 respectively

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 Δt is the temperature change plugging
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 the heart of all calorimetry problems
 and calculations because the heat
 released or absorbed at constant
 pressure is equal to Δh the relationship
 between heat and Δh_{rxn} is $\Delta h_{rxn} q_{rxn}$
 $q_{calorimeter} mcs\Delta t$ 7 3 14 7 3 14 Δh_{rxn}
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 1 3 i have 18 7 grams of an unknown
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 of the calorimeter and its contents
 changes from 19 77 c to 33 90 c
 calculate e and h for the following
 reaction as written 2 nh 3 g 3 o 2 g 2 ch

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ms t h 5.650 grams H_2O 4.184 J/g °C 55.4 °C 1310 kJ

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