

# A Good Traveler Lao Tzu

**Lingsheng Yao**

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both compound types acids ionic

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ions ionic compounds with uncommon

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web jul 25 2022 molecular compounds

can form compounds with different

ratios of their elements so prefixes are

used to specify the numbers of atoms of

each element in a molecule of the

compound examples include sf<sub>6</sub> sulfur

hexafluoride and n<sub>2</sub>o<sub>4</sub> dinitrogen

tetroxide

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web the name of a binary compound

containing monatomic ions consists of

the name of the cation the name of the

metal followed by the name of the anion

the name of the nonmetallic element

with its ending replaced by the suffix

ide some examples are given in table 2

6 names of some ionic compounds table

2 6 compounds containing

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and empirical formulas of the following

compounds a b c d answer a molecular

c 4 h 8 empirical ch 2 answer b

molecular c 4 h 6 empirical c 2 h 3

answer c molecular h 2 si 2 cl 4

empirical hsiCl<sub>2</sub> answer d molecular h

3 po 4 empirical h 3 po 4 click here for

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web bismuth has a molar mass of one

mole of the compound would result in

two moles of bismuth a total of 417 96g

tellurium has a molar mass of one mole

of the compound would result in three

moles of tellurium a total of 382 8g add

the mass of bismuth and the mass of

tellurium per mole to find the total

molar mass divide the mass of bismuth

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compounds atoms held together by the

sharing of electrons in covalent bonds

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web in molecular compounds the atom

binds each other through covalent

bonds in salts it is held together with

ionic bonds these are the two types of bonds out of which every compound is made of 2 example of compounds example of compounds includes water  $\text{H}_2\text{O}$  hydrogen peroxide  $\text{H}_2\text{O}_2$  etc

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web molecular compounds are inorganic compounds that take the

form of discrete molecules examples include such familiar substances as water  $\text{H}_2\text{O}$   $\text{H}_2\text{O}$  and carbon dioxide  $\text{CO}_2$   $\text{CO}_2$  these compounds are very different from ionic compounds like sodium chloride  $\text{NaCl}$   $\text{NaCl}$

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