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curve is a plot of the ph of a solution as
a function of the amount of base usually
in ml added to an acidic solution or as
the amount of acid added to a base
solution the base is added to

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bases the shape of the titration curve
for a weak acid or base depends on the
 pK_a or pK_b of the weak acid or base
being titrated a solution ph as a
function of the volume of 1 00 M
NaOH added to 10 00 ml of 1 00 M
solutions of weak acids with the
indicated pK_a values

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theory procedure self evaluation
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amount of substance in a solution of
unknown concentration using various
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word titration comes from the latin
word titulus which means inscription or
title

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web where the salt is the conjugate
base of the acid the four parts of the
titration curve are described below and
you should look to the appropriate text
section to see how they are treated
pure acid 0 ml of base is added section
17 3 2 1

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web solution a titrant volume 0 ml the
solution ph is due to the acid ionization
of HCl because this is a strong acid the
ionization is complete and the
hydronium ion molarity is 0 100 M the
ph of the solution is then $ph = -\log 0 100$
1 000 $ph = -\log 0 100 = 1 000$ b titrant volume
12 50 ml

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analytical procedure used to determine
the accurate concentration of a sample
by reacting it with a standard solution
one type of titration uses a
neutralization reaction in which an acid
and a base react to produce a salt and
water

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web at the equivalence point in an acid
base titration moles of base moles of
acid and the solution only contains salt
and water acid base titrations are
monitored by the change of ph as
titration progresses

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example is based upon henderson
hasselbalch equation the equation gives
relation between ph and pH_a of weak
acid the equation is $ph = pK_a + \log \frac{[salt]}{[acid]}$
at equivalence point all a view the
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web unlike strong acids or bases the
shape of the titration curve for a weak
acid or base depends on the pK_a or pK_b

b of the weak acid or base being titrated a solution pH as a function of the volume of 1.00 M NaOH added to 10.00 mL of 1.00 M solutions of weak acids with the indicated pK_a values

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are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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