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web $m \times c \times \Delta t$ equation 1 where m is the total mass of the
solution solute plus solvent c is the specific heat of the
solution and Δt is the observed temperature change the
specific heat of the solution is generally assumed to be the
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web the heat of reaction ΔH_{soln} is written after the products
in units of kJ mol rxn for example sodium hydroxide dissolves
exothermically releasing 44.2 kilojoules per mole dissolved
this process is represented as $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$ $\Delta H_{\text{soln}} = -44.2$
kJ mol rxn write equations to similarly represent the

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web hand warmers are usually placed inside the gloves to
keep your hands warm the heat keeping your hands warm is
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considered dangerous do not touch taste or smell any
chemicals unless specifically instructed to do so 2 never
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dissolves in water and produces a large temperature change
in the form of heat in this laboratory experiment the heat of
solution will be calculated from the data for six salts in order

to

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution *hand warmer chemistry easy chemical hot packs science - Apr 30 2023*

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