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 g 5 CO_2 g I_2 g a 80 0 grams of iodine v

oxide i_2O_5 reacts with 28 0 grams of
 carbon monoxide co determine the
 mass of iodine I_2 which could be
 produced b if in the above situation
 only 0 160 moles of iodine I_2 was
 produced i what mass of iodine was
 produced

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 contains 0 mol KO_2 and 0 mol h 2 o
 what is the limiting reactant

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 quantities of two or more reactants are
 given in a stoichiometric problem you
 must identify the click the card to flip
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web limiting reactant reactant present
 in an amount lower than required by
 the reaction stoichiometry thus limiting
 the amount of product generated
 percent yield measure of the efficiency

of a reaction expressed as a percentage
 of the theoretical yield

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 following reaction $\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$
 agno 3 \rightarrow 2 text ag text cu no 3
 text 2 cu 2agno3 2ag cu no3 2

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 percentage yield 1 for the balanced
 equation shown below if the reaction of
 90 6 grams of co produces 36 7 grams
 of C_3H_8 what is the percent yield $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8 3 mole H_2S and
 multiply it with 2 mole H_2O over 2 mole
 H_2S from the equation $8 3 \text{ mole } \text{H}_2\text{S} \times 2$
 $\text{mole } \text{H}_2\text{O} = 2 \text{ mole } \text{H}_2\text{S} \times 18\text{g } \text{H}_2\text{O} = 1 \text{ mole}$
 $\text{H}_2\text{O} = 149 4\text{g } \text{H}_2\text{O}$ is your theoretical yield
 so percent yield $137 1\text{g } \text{H}_2\text{O} / 149 4\text{g } \text{H}_2\text{O} \times 100 = 91 77\%$
 problems 6 8

are similar to problem 5

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80 is anticipated

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h2

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g co 1 mol co

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