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beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0 100 m the ph of the solution is then $\text{ph} = -\log 0 100 = 1 000$ ph $\log 0 100 = 1 000$ b titrant volume 12 50 ml

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web acid and base titrations lab report chm 114 jx abstract this goal was to give us experience finding the standardization of through the use of a primary standard in this experiment we will be using naoh and hcl as well as khp

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web question lab report acid base titrations purpose of the experiment in this experiment we will titrate an acid with a base to determine the concentration of an acid with a base to determine the concentration of an acidic sample