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 $2 \text{ CO}_2 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + \frac{5}{2} \text{ O}_2 \text{ g}$ $\Delta h = 2 \text{ h}_2 \text{ g} + \text{C}_2\text{H}_6 \text{ g}$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O g} \rightarrow \text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g}$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2 \text{ O g}$ $\Delta h = 283 \text{ kJ}$
 $\text{g} + 1 \text{ C}_2\text{H}_6 \text{ g}$

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 $\text{C}_2\text{H}_4 \text{ g} + 3 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 2 \text{ H}_2 \text{ O l}$ $\Delta h = 1411 \text{ kJ}$
 $\text{C}_2\text{H}_6 \text{ g} + 7 \text{ O}_2 \text{ g} \rightarrow 2 \text{ CO}_2 \text{ g} + 3 \text{ H}_2 \text{ O l}$ $\Delta h = 1560 \text{ kJ}$
 $\text{H}_2 \text{ g} + \frac{1}{2} \text{ O}_2 \text{ g} \rightarrow \text{H}_2 \text{ O l}$ $\Delta h = 285 \text{ kJ}$
 mole

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o2 g 2 co2 g 2 h2o l h 1411 kj c2h6 g 3½ o2 g 2 co2 g 3 h2o l h 1560 kj

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web thus for reaction 9 2 in which 1 mole of h3o and 1 mole of oh combine to form 2 moles of h2o h 55 8 kj as shown note that the enthalpy change h hproducts hreactants is positive if heat is absorbed that is if hproducts

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