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web oct 28 2019 solution use this equation $q = \text{specific heat} \times m \times \Delta t$ where q is heat flow m is mass in grams and Δt is the temperature change plugging in the values given in the problem you get $q_{\text{water}} = 4.18 \text{ j/g} \times 110 \text{ g} \times 26.6 \text{ c} = 1250 \text{ c} \times q_{\text{water}} = 550 \text{ j}$ $\Delta h_{\text{q water}} = 550 \text{ j}$

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t q heat m mass qrxn qsur t tf ti name per date c specific
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released or absorbed at constant pressure is equal to Δh the
relationship between heat and Δh_{rxn} is $\Delta h_{rxn} q_{rxn}$
 $q_{calorimeter} mcs\Delta t$ 7 3 14 7 3 14 $\Delta h_{rxn} q_{rxn} q_{calorimeter} mcs\Delta t$

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questions 1 if i have 125 grams of water at a temperature of
25 0 c and heat it to 145 0 c what will the change in enthalpy
be 2 sketch the heating curve temperature vs time for the
process in 1 3 i have 18 7 grams of an unknown liquid
chemical at a temperature of 35 0 c

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experiment 25 0 ml of 1 00 m hcl at 25 0 c is added to 25 0
ml of 1 00 m naoh at 25 0 c in a foam cup calorimeter a
reaction occurs and the temperature rises to 32 0 c calculate
the enthalpy change in kj for this reaction assume the
densities of the solutions are 1 00 g ml and that

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chemical reactions consider a simple example that illustrates
the core idea behind calorimetry suppose we initially have a
high temperature substance such as a hot piece of metal m
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answer the initial temperature of the copper

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reaction as written 2 nh 3 g 3 o 2 g 2 ch

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